

CHEM 133 Exam 1
Mar. 3, 2015
KEY

Equations and constants that you may find useful:

Equation for voltage across resistor in RC circuit as a function of time for a step change of ΔV_{in} at time $t = 0$:

$$V_R = \Delta V_{in} e^{-t/RC}$$

Standard deviation:

$$S = \sqrt{\frac{\sum(x_i - \bar{x})^2}{n-1}}$$

Digitization equation for n bit digitizer: decimal # = $(V_{meas} - V_{min}) \cdot 2^n / V_{range}$

3σ limit of detection: $C_{LoD} = 3\sigma/m$ where m = slope of response vs. conc. (C) plot and σ = standard deviation of blank or noise data.

Noise Equations: 1) Thermal noise: $V_{noise}(rms) = (4kTR\Delta f)^{1/2}$

2) Shot Noise: $I_{noise}(rms) = (2eI\Delta f)^{1/2}$

k = Boltzmann's constant = $1.38 \times 10^{-23} \text{ V}^2 \text{ s } \Omega^{-1} \text{ K}^{-1}$, T = temperature (K), R = resistance, Δf = bandwidth, I = current, and e = elementary charge = 1.60×10^{-19} coulombs.

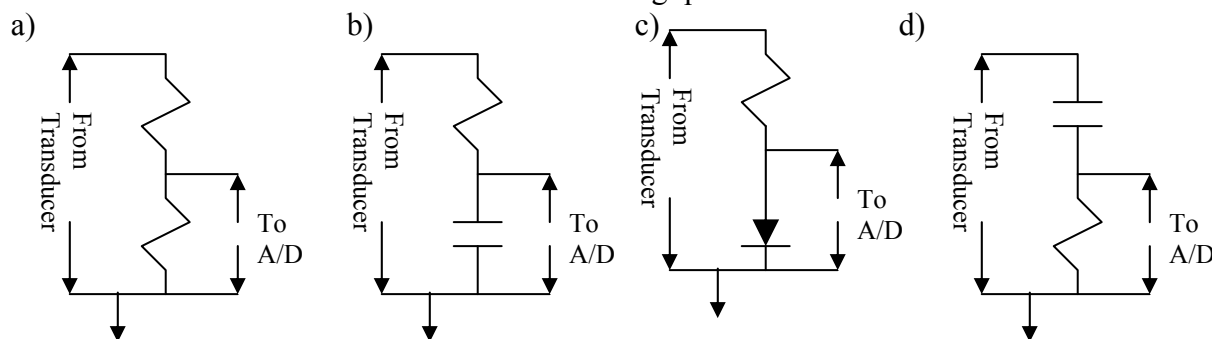
F = Faraday's constant = 96,500 C/mol e

SHORT ANSWER SECTION: (Each question worth 4 points)

1. Chemist A is investigating the presence of highly toxic mold compounds in processed peanuts, where very small concentrations of mold compounds can have significant effects. Chemist B is measuring the types of oils in the peanuts. Oils are a major fraction of peanut mass. Compared to Chemist B, Chemist A needs excellent:

- a) accuracy b) precision **c) sensitivity** d) sample throughput

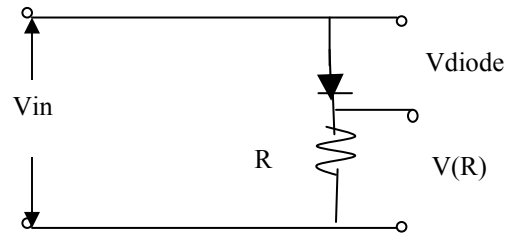
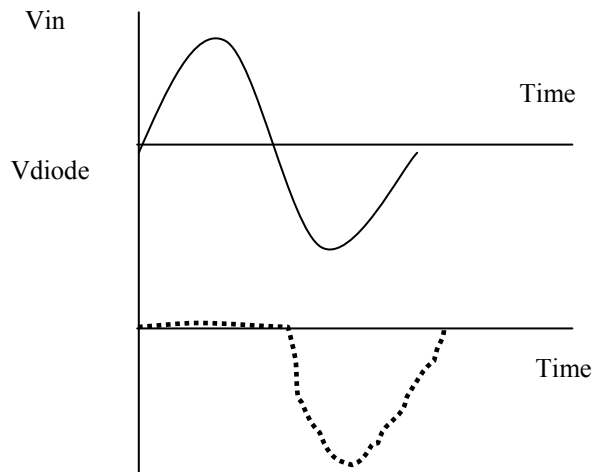
2/3. Given the circuits below answer the following questions



2. Which of the above circuits is best to linearly decrease a transducer voltage output (mainly DC signal) so that it is in the range of an analog to digital input (A/D input)? **a**

3. Which above circuit is best for removing high frequency noise? (e.g. a low pass filter) **b**

2. Assuming the diode in the figure to the right behaves ideally and V_{in} is as shown to the left, sketch V_R and V_{diode} as a function of time. (6 pts)

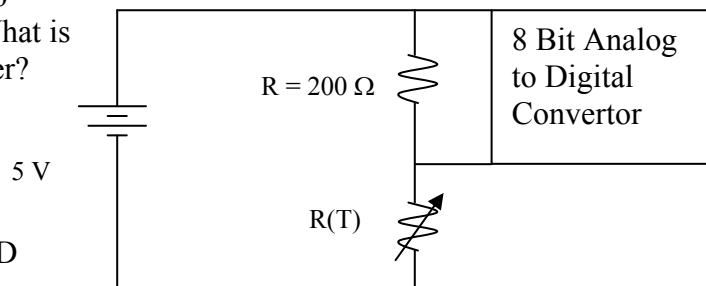


Note: when $V_{in} > 0$, $R_{diode} = 0$, so $V_{diode} = 0$ and $V_R = V_{in}$
 when $V_{in} < 0$, $R_{diode} = \text{large}$, so $V_{diode} = V_{in}$ and $V_R = 0$

3. The circuit shown below is used to measure temperature with a thermistor $R(T)$. The resistance of the thermistor, $R(T)$, is equal to $192 \text{ k}\Omega \cdot \text{K}/T$ (where T is the absolute temperature in units K), and the input range of the analog to digital convertor is 0 to 2 V. (26 pts)

a) The binary output of the analog to digital convertor is 10011011. What is the corresponding decimal number?

$$1 + 2 + 0 + 8 + 16 + 0 + 0 + 128 = 155$$



b) What is the voltage read by the A/D convertor?

$$155 = (V_{meas} - V_{min}) \cdot 2^n / V_{range} = (V_{meas} - 0) \cdot 2^8 / (2 - 0)$$

$$V_{meas} = 155 \cdot 2V / 256 = \mathbf{1.211 \text{ V}}$$

b) What is the value of $R(T)$ in Ω based on the calculation from part a)?

$$R(T) = V_{R(T)} / I \text{ and } 5.00 \text{ V} = V_{R(T)} + V_R \text{ or } V_{R(T)} = 5.00 - 1.211 \text{ V} = 3.789 \text{ V}$$

$$\text{Also, } I = V_R / R = 1.211 \text{ V} / 200 \Omega = 0.006055 \text{ A}$$

$$R(T) = 3.789 / 0.006055 = 625.8 = \mathbf{626 \Omega}$$

c) What temperature (in K) does this resistance correspond to?

$$626 = 192,000 \Omega \text{ K}/T \text{ or } T = 192,000 / 626 = 306.8 \text{ K} = \mathbf{307 \text{ K}}$$

Bonus problem – with 3: What is the average error in temperature from digitization at 298K? (3 pts)

The average voltage error from digitization = $0.5(1/2^8)2V = 0.00391 V$

Now we need to translate this to an error in T.

$5.00 V = I(R + 192000 \Omega K / T)$ and $I = V_{meas}/R$ so $5.00 V = (V_{meas}/R)(R + 192000 \Omega K / T)$

$(5.00 V)(200 \Omega)/V_{meas} = 200 \Omega + 192000 \Omega K / T$

At this point we can 1) calculate V_{meas} at 298 and then add 0.00391 V to V_{meas} , recalculate T (e.g. as $T + \Delta T$) and take the difference or 2) solve the above equation for T and differentiate T with respect to V, and do a calculus based propagation of uncertainty calculation

From method 1) $V_{meas} = 1000/(200 \Omega + 192000 \Omega K / 298K) = 1.18442 V$ and $V + \Delta V = 1.1883 V$

$1000/1.18833 - 200 = 192000 \Omega K / (T + \Delta T)$, where $T + \Delta T = 299.3 K$ and $\Delta T = 1.3 K$

In method 2) $dT/dV_{meas} = 1.92 \times 10^8 / (10^6 - 4 \times 10^5 V_{meas} + 4 \times 10^4 V_{meas}^2) = 330.6$

And $dT = 330.6(dV_{meas}) = 1.3 K$

4. A mass spectrometer is used to analyze a sample for a 2.0 s run. A peak at 95 mass units is observed with a signal to noise of 26.3. The number of scans performed is proportional to the time. It is suspected that the compound has one S atom in it which would give a 97 peak with 4.5% of the mass 95 signal. How long of a run is needed to see the 97 peak with a signal to noise ratio of 3? Assume the noise is random. (12 pts)

Peak at 95 has S/N of 26.3. If the noise is the same for each peak (not specified but assumed),

Peak at 97 should have a signal of 0.045(signal for 95 peak). With noise assumption,

$S/N(97 \text{ peak}) = 26.3(0.045) = 1.18$

We know that $[S/N(97 \text{ peak desired})]/[S/N(97 \text{ peak } 2 \text{ s run})] = [n(\text{desired})/n(2 \text{ s run})]^{0.5}$

and also that $n = kt$ so that $n(\text{desired})/n(2 \text{ s run}) = t(\text{desired})/2 \text{ s}$

$3/1.18 = [t(\text{desired})/2 \text{ s}]^{0.5}$ or $t(\text{desired}) = (3/1.18)^2(2 \text{ s}) = 13 \text{ s}$

5. A chrome plating bath uses CrO_4^{2-} in an acidic solution, which is reduced to Cr(s) in the reaction. If a muffler requires 38 g of Cr (atomic weight = 52.00 g/mol), and a constant current of 30.0 A is used. (16 pts)

a) Calculate the time in hours needed for the chrome plating.

We can start by relating the mass of Cr to the moles of electrons. Method 1) Cr in CrO_4^{2-} has an oxidation state of +6 vs. 0 for Cr in Cr(s), so we need 6 electrons. Method 2) create balanced half reaction: $\text{CrO}_4^{2-} + 8\text{H}^+ + 6\text{e}^- \rightarrow \text{Cr}(s) + 4\text{H}_2\text{O}(l)$

$Q = \text{charge} = n_e F = 6n_{\text{Cr}} F = (6 \text{ mol } e/\text{mol Cr})(38 \text{ g Cr})(1 \text{ mol Cr}/52.00 \text{ g Cr})(96,500 \text{ C/mol } e)$

$Q = 423,115 \text{ C} = I \cdot t$ where $I = \text{current}$ and $t = \text{time}$

$t = 423,115 \text{ C}/(30.0 \text{ C/s}) = (14100 \text{ s})(1 \text{ min}/60 \text{ s})(1 \text{ hr}/60 \text{ min}) = 3.9 \text{ hr}$

- b) If the plating requires an applied voltage of 2.85 V, what is the power consumption of the chrome plating operation?

$P = IV = (30.0 \text{ A})(2.85 \text{ V}) = 85.5 \text{ W}$