

CHEM 133 Exam 1

Mar. 7, 2017

KEY

Equations and constants that you may find useful:

Equation for voltage across resistor in RC circuit as a function of time for a step change of ΔV_{in} at time $t = 0$:

$$V_R = \Delta V_{in} e^{-t/RC}$$

Standard deviation:

$$S = \sqrt{\frac{\sum (x_i - \bar{x})^2}{n-1}}$$

Digitization equation for n bit digitizer: decimal # = $(V_{meas} - V_{min}) \cdot 2^n / V_{range}$

Noise Equations: 1) Thermal noise: $V_{noise(rms)} = (4kTRB)^{1/2}$

2) Shot Noise: $I_{noise(rms)} = (2eIB)^{1/2}$

k = Boltzmann's constant = $1.38 \times 10^{-23} \text{ V}^2 \text{ s } \Omega^{-1} \text{ K}^{-1}$, T = temperature (K), R = resistance, B = bandwidth, I = current, and e = elementary charge = $1.60 \times 10^{-19} \text{ C}$. C = coulombs

F = Faraday's constant = 96,500 C/mol e

Nernst Equation:

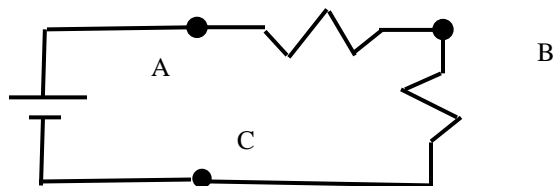
$$E = E^\circ - \frac{RT}{nF} \ln Q = E^\circ - \frac{2.303RT}{nF} \log Q = E^\circ - \frac{0.05916}{n} \log Q \text{ (T = 298 K)}$$

R = Universal Gas Constant = $8.314 \text{ J mol}^{-1} \text{ K}^{-1}$; n = moles e^-

Units: 1 A = 1 C/s; 1 J = 1 C · V; 1 W = 1 V · A, A = amps, V = volts, W = watts, J = joules

SHORT ANSWER SECTION: (Each question worth 4 points)

1. Adding a resistor where will cause an increase in current from the power supply?



a) inserted at point A

b) inserted at point B

c) connected between points A and C

d) any addition of a resistor will decrease current

2. A fluoride selective electrode has a useful output range of 0.6 to 2.1 V and is being recorded by a digitizer with an input voltage range of 0 to 1 V. Noise is not a major issue. What type of analog signal processing is needed?

a) RC high pass filter

b) RC low pass filter

c) a voltage divider

d) a voltage amplifier

3. List a transducer for temperature measurement and the electrical output signal (e.g. voltage, current or resistance). Transducer a) thermistor, b) PRT, c) thermocouple Output signal a/b) R, c) V__

4. A 10 bit analog to digital convertor allows input voltages from -500 mV to +500 mV. What is the average error in the signal read (in mV) following digitization?

- a) 0.24 mV **b) 0.49 mV** c) 50 mV d) 100 mV

ave error = $\frac{1}{2}(\text{smallest bit}) = (1/2)[500 \text{ mV} - (-500 \text{ mV})](1/2^{10})$

5. What type of noise is associated with slow drifts in signal response and can be reduced by modulating part of an instrument (to perform measurements at higher frequency)?

- a) 1/f noise** b) shot noise c) thermal noise d) Johnson noise

6. Which of the following quantities is most important in determining the amount of charge available from a battery:

- a) E° values for the reaction b) Q, the reaction quotient
c) the moles of the limiting reagent d) Zeta, the kinematic supply coefficient

7. In an electrolytic cell,

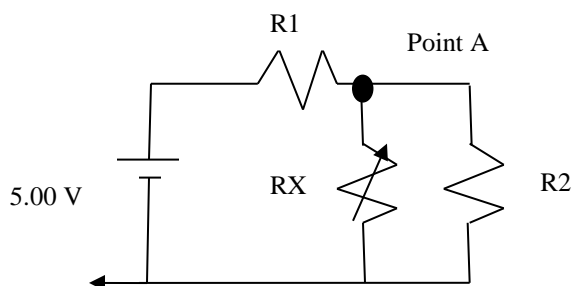
- a) chemical potential is used to perform electrical work
 b) reduction occurs at the anode
 c) a reference electrode is required
d) an external potential is used to carry out unfavorable reactions

SECTION II. PROBLEM SECTION. Show work – use the back side if needed

1. In the circuit to the right, if $R_1 = 220. \Omega$ and $R_2 = 370. \Omega$, and V_A is measured as 2.31 V, calculate the following:

- a) The current output at the power supply (6 pts)

$$\begin{aligned} I(\text{power supply}) &= I_{R1} \\ &= V_{R1}/I_{R1} = (5.00 \text{ V} - 2.31 \text{ V})/220 \Omega \\ I(\text{power supply}) &= \mathbf{0.0122 \text{ A}} \end{aligned}$$



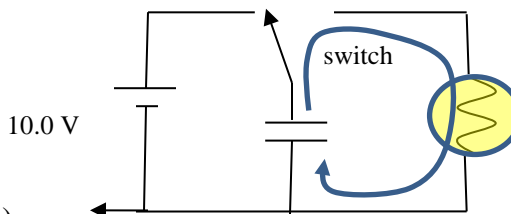
- b) The current across resistor RX (6 pts)

$$\begin{aligned} I_{R1} &= I_{R2} + I_{RX} \text{ or } I_{RX} = I_{R1} - I_{R2} \text{ and } I_{R2} = V_A/R_2 = 2.31 \text{ V}/370 \Omega = 0.00624 \text{ A} \\ I_{RX} &= 0.0122 \text{ A} - 0.00624 \text{ A} = \mathbf{0.0060 \text{ A}} \end{aligned}$$

- c) The resistance of RX. (4 pts)

$$RX = V_A/I_{RX} = 2.31 \text{ V}/0.0060 \text{ A} = \mathbf{390 \Omega}$$

2. A capacitor ($C = 50.0 \mu\text{F}$) is charged in the circuit to the right and then discharged through a light that can be treated as a resistor ($R = 40.0 \Omega$). If the capacitor is charged to 10.0 V and light will be emitted until the voltage drops to 3.0 V , how long will the light be on for after flipping the switch? (8 pts)



Note: switch will move from left side to right side to start capacitor discharge.

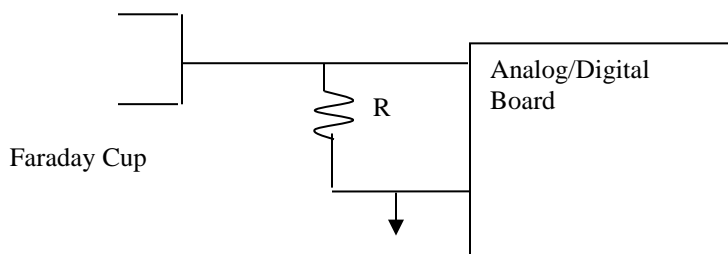
After flipping switch, 10 V in charged capacitor gets dissipated as it goes through the lamp as shown in the loop above. From Kirchhoff's potential law, we can write,

$$V_C = V_R = (10.0\text{V})e^{-t/RC} = 3.0\text{V} \text{ or } 0.30 = e^{-t/RC} \text{ or } \ln(0.30) = -t/RC$$

$$t = -\ln(0.30)(40.0 \Omega)(50.0 \times 10^{-6} \text{ F}) = \mathbf{0.0024 \text{ s}}$$

3. In the circuit shown below, a Faraday cup in a flame ionization detector (FID) is used to measure the current positive ions produce in the flame as they hit the cup. The A/D board records the voltage drop across the resistor with resistance $= 135,000 \Omega$ so that the Faraday cup current can be calculated. The FID current is proportional to flux (nmol C reaching the FID per second). The A/D board has 10 bits with a range of 0.000 to 0.500 V .

Answer the following questions below:



When 5.0 nmol C/s reaches the FID, the average binary number recorded by the A/D board is: 0101001110.

a) What is the corresponding decimal number? (5 pts)

we can get the decimal number by writing the 2 sequence and cancelling out any with 0s.

$$\mathbf{512 + 256 + 128 + 64 + 32 + 16 + 8 + 4 + 2 + 1 = 334}$$

b) What voltage corresponds to A/D reading? (6 pts)

$$\text{decimal \#} = (V_{\text{meas}} - V_{\text{min}}) \cdot 2^n / V_{\text{range}} \text{ (see front page)}$$

$$334 = (V - 0)(1024)/0.500 \text{ V}$$

$$\text{Or } V = (334)(0.500 \text{ V})/1024 = \mathbf{0.163 \text{ V}}$$

c) What is the current coming from the Faraday cup? (assume zero current through A/D board) (5 pts)

$$I = V/R = 0.163 \text{ V}/(135,000 \Omega) = 1.21 \times 10^{-6} \text{ A} = \mathbf{1.21 \mu\text{A}}$$

3 – cont. – Bonus) Give the maximum detectable flux (in nmol/s). (3 pts)

Flux is proportional to current and to voltage so $F_{\max}/F = V_{\max}/V$

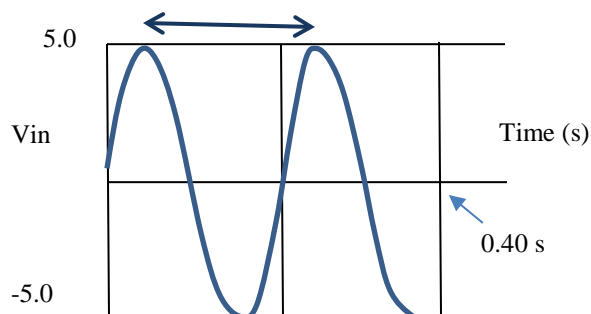
or $F_{\max}/(5.0 \text{ nmol/s}) = (0.500 \text{ V}/0.163 \text{ V})$ or $F_{\max} = (5.0 \text{ nmol/s})(3.07) = 15 \text{ nmol/s}$

4. Given the circuit and V vs. time plot to the right answer the following questions. (8 pts)

a) What is the frequency of V_{in} ?

$f = 1/(\text{time period})$ (see arrow on plot)

$f = 1/0.2 \text{ s} = 5 \text{ Hz}$



b) If the plot above represents a perfect sine wave and extends infinitely along the time axes, describe what a Fourier transform of it would look like.

If a sine wave, a single line at 5 Hz would show up.

5. A compound is being analyzed by mass spectrometry and may have one or two sulfurs. Having one sulfur gives rise to an $M + 2$ isotope peak (from having an ^{34}S in place of a ^{32}S) about 4.5% as large as the M peak. When data is averaged for 50 scans, the signal to noise ratio for the M peak is 25. How many scans should the mass spectrometer run so that the signal to noise ratio in the $M + 2$ peak is 5 assuming 1 S present? (12 pts)

For M peak, $S/N = 25$. For $M + 2$ peak, signal is 4.5% as large so $S/N = (0.045)(25) = 1.125$

The desired S/N (for $M + 2$ peak) is 5 and the S/N is proportional to $n^{0.5}$

Or $(S/N)_n/(S/N)_{n=50} = (n/50)^{0.5}$ or $5/1.125 = (n/50)^{0.5}$ or $(5/1.125)^2 = n/50$

$n = (19.8)(50) = 988 = 990$ (2 sig fig)

6. A chemist is testing an electrode for measuring Cd^{2+} in solution by using the following cell: $\text{Cd(s)}|\text{Cd}^{2+}(\text{aq}, 0.0051 \text{ M})||\text{KCl(aq}, 1.0 \text{ M})|\text{AgCl(s)}|\text{Ag(s)}$. Given the E° for the following reactions, calculate the cell potential (remember in the cell notation that the left half is for oxidation). (12 pts)

Standard reduction potentials: $\text{Cd}^{2+}(\text{aq}) + 2\text{e}^- \leftrightarrow \text{Cd(s)}$ $E^\circ = -0.380 \text{ V}$

$\text{AgCl(s)} + \text{e}^- \leftrightarrow \text{Ag(s)} + \text{Cl}^-(\text{aq})$ $E^\circ = 0.222 \text{ V}$

$E_{\text{cell}} = E_{\text{cathode}} - E_{\text{anode}}$ and $E_{\text{cathode}} = E^\circ - (0.05916/1)\log[\text{Cl}^-] = 0.222 \text{ V} - 0$

$E_{\text{anode}} = E^\circ - (0.05916/2)\log 1/[\text{Cd}^{2+}]$

$E_{\text{cell}} = 0.222 \text{ V} - [-0.380 \text{ V} - (0.05916/2)\log(1/0.0051)] = 0.602 + 0.068 = 0.670 \text{ V}$