1. (a) Al(NO$_3$)$_3$  (b) (NH$_4$)$_2$O  (c) dinitrogen tetraoxide  (d) copper (II) cyanide

2. D
3. C
4. C
5. A
6. 13.5 g/cm$^3$ = 4 pts

1. No partial credit

2. All work must be shown
   - 1 pt for all incorrect answers
   - ½ pt for each missing unit
   - No credit if the set up is wrong.

\[
d = \frac{\frac{\text{mass}}{\text{volume}}}{\text{density}} = \frac{0.39 \text{ kg} \times \frac{10.5}{\text{kg}}}{10.0 \text{ oz} \times \frac{29.6 \text{ mL}}{102 \text{ oz}} \times \frac{1 \text{ cm}^3}{1 \text{ mL}}} = 13.5 \text{ g/cm}^3
\]

Ans. Must have correct units, i.e. no units, deduct ½ pt.