By definition:

\[\% \text{ (W/V)}\]

Percent of weight of solution in the total volume of solution. Percent here is the number of grams of solute in 100 mL of solution.

**Example:**

A 4\% (W/V) NaCl solution is 4 g of NaCl in 100 mL of solution.

Your question:

How to make one liter 0.8\% solution from 0.9\% solution

1) We know that:

\[
0.9\% = \frac{0.9}{100} = 0.009 \quad \text{(mg/mL)}
\]

\[
0.8\% = \frac{0.8}{100} = 0.008 \quad \text{(mg/mL)}
\]

1 L = 1000 mL

2) Next use the following relationship as:

\[
(0.009) (X) = (0.008) (1000)
\]

Solve it for the (X):

\[X = 889 \text{ mL}\]

It means add 889 mL of the 0.9\% solution to a one liter volumetric flask and add DI water to its mark.

Or, it means add (1000- 889=111 mL) of DI water to 889 mL of 0.9\% solution.