

16) $A = \epsilon bc$

a. $c = \frac{A}{\epsilon b} = \frac{0.427}{(6130 \text{ M}^{-1} \text{ cm}^{-1})(1.000 \text{ cm})} = \boxed{6.97 \times 10^{-5} \text{ M}}$

b. $\frac{6.97 \times 10^{-5} \text{ M}}{(1/10)} = \boxed{6.97 \times 10^{-4} \text{ M}}$

c. $5.00 \text{ mL} \left| \frac{6.97 \times 10^{-4} \text{ mol}}{1000 \text{ mL}} \right| \left| \frac{292.16 \text{ g}}{1 \text{ mol}} \right| \left| \frac{1000 \text{ mg}}{1 \text{ g}} \right| = \boxed{1.02 \text{ mg}}$

18) $A = \epsilon bc$

a. $\epsilon = \frac{A}{bc} = \frac{(0.267 - 0.019)}{(1.00 \text{ cm})(3.15 \times 10^{-6} \text{ M})} = \boxed{78730 \text{ M}^{-1} \text{ cm}^{-1}}$

b. $c = \frac{A}{\epsilon b} = \frac{(0.175 - 0.019)}{(78730 \text{ M}^{-1} \text{ cm}^{-1})(1.00 \text{ cm})} = \boxed{1.98 \times 10^{-6} \text{ M}}$

19) a. Absorbance due to known amount of NO_2^-
 $= 0.967 - 0.622 = 0.345$

concentration of known $\text{NO}_2^- = 7.50 \times 10^{-3} \text{ M} (10.08 \times 10^{-6} \text{ L})$

$\epsilon = \frac{A}{bc} = \frac{0.345}{(5.00 \text{ cm})(1.39 \times 10^{-6} \text{ M})} = \frac{7.5 \times 10^{-8} \text{ mol}}{0.054 \text{ L}} = 1.39 \times 10^{-6} \text{ M}$
 $= \boxed{49680 \text{ M}^{-1} \text{ cm}^{-1}}$

b. $\frac{7.50 \times 10^{-8} \text{ mol}}{x \text{ mol}} = \frac{0.345}{0.469} \Rightarrow x = 1.02 \times 10^{-7} \text{ mol NO}_2^-$

$1.02 \times 10^{-7} \text{ mol NO}_2^- \left| \frac{46 \text{ g}}{1 \text{ mol}} \right| \left| \frac{10^6 \mu\text{g}}{1 \text{ g}} \right| = \boxed{4.69 \mu\text{g NO}_2^-}$

24 Fluorescence is emission of a photon with an electronic transition from an excited singlet state to the ground singlet state. The lifetime of the excited singlet state is very short.

Phosphorescence is emission of a photon with an electronic transition from a triplet state to a singlet state. The lifetime of the excited triplet state is relatively long.

25 Luminescence results after a molecule absorbs light. Chemiluminescence results from the product of a chemical reaction that is in the excited state.

28 For an excitation spectrum the λ of the spectrometer's light source is varied and light of a specific emission λ is monitored. For an emission spectrum, the excitation λ is held constant and the measured emission λ is varied. The excitation spectrum resembles an absorption spectrum because emission is proportional to absorption.