

Useful Information for Gas Law Calculations**Pressure Measurement Units:**

<u>Unit:</u>	<u>Conversion to atm:</u>
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Atmosphere (atm)	1
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760 mm Hg	= 1 atm
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760 torr	= 1 atm
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*1 mm Hg = 1 torr*

101,325 Pa*	= 1 atm
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1.01325 bar**	= 1 atm
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14.696 lb/in <sup>2</sup>	= 1 atm
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\*Pa (N/m<sup>2</sup>) are often written as kPa            (1 kPa = 10<sup>3</sup> Pa)

\*\*bar are often written as mbar            (1 mbar = 10<sup>-3</sup> bar)

**Gas Constant:**

$$R = 0.08206 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$$

**Temperature:**

$$K = 273.15 + ^\circ\text{C}$$

**Standard Temperature and Pressure:**

Temperature = 0 °C            Pressure = 1 atm