

A weak electrolyte exists predominantly as \_\_\_\_\_ in solution.

- a) molecules
- b) ions
- c) atoms
- d) an isotope
- e) electrons

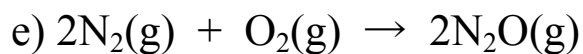
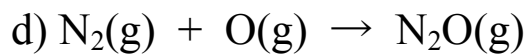
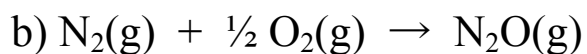
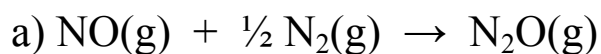
What is the predominant intermolecular force in  $\text{CH}_4$ ?

- a) dipole-dipole attraction
- b) ion-dipole attraction
- c) hydrogen-bonding
- d) London-dispersion forces
- e) ionic bonding

How many cubic centimeters ( $\text{cm}^3$ ) are in  $15.2 \text{ in}^3$ ?

- a. 38.6      b. 249      c. 155      d. 5.98      e. 0.27

Which of the following chemical equations corresponds to the standard molar enthalpy of formation of  $\text{N}_2\text{O}$ ?



Which one of the following pictures best represents what an aqueous solution of  $\text{Na}_2\text{SO}_4$  looks like on the molecular level?

*Water molecules are not shown for clarity.*

