

ELECTROLYTES:

What are they?

Three types of electrolytes:

Non-electrolytes

Molecular compounds

Do NOT form ions in water

Weak electrolytes

Weak acids

Insoluble salts (based on solubility rules)

Form very VERY few ions in water

Strong electrolytes:

Strong acids

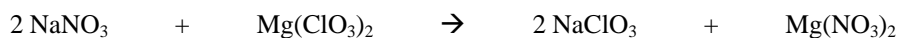
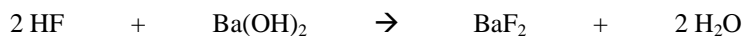
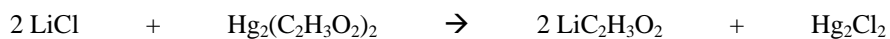
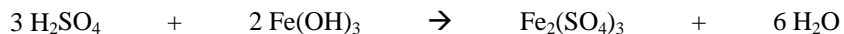
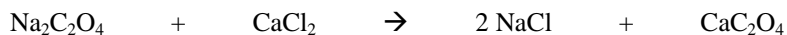
Soluble salts (based on solubility rules)

Completely break apart into ions when put in water

Called *ionization* when talking about acids

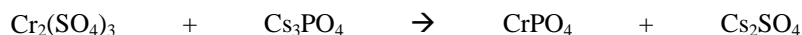
Called *disassociation* when talking about salts

SOLUBILITY RULE PRACTICE/ELECTROLYTE PRACTICE



NET IONIC EQUATIONS:

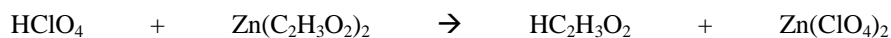
- 1) Start with a balanced chemical equation (including all states of matter)



Break ALL strong electrolytes into ions. Leave non-electrolytes and weak electrolytes alone!

Cross out everything that is EXACTLY the same on both sides of the equation.

- 2) Start with a balanced chemical equation (including all states of matter)



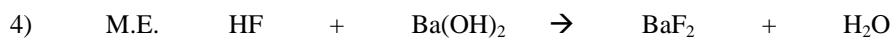
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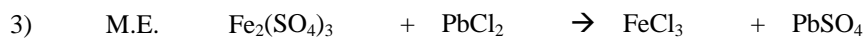
I.E.

N.I.E.



I.E.

N.I.E.



I.E.

N.I.E.



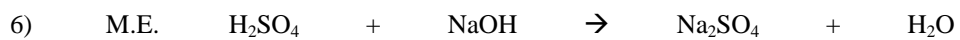
I.E.

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