

Solubility:

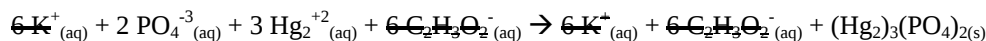
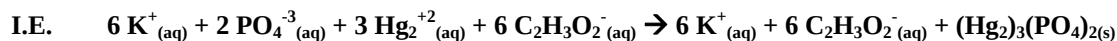
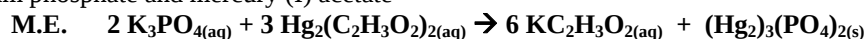
1)	$(\text{NH}_4)_2\text{S}_{(\text{aq})}$	rule #1
2)	$\text{BaO}_{(\text{aq})}$	rule #6
3)	$\text{NaCO}_3_{(\text{aq})}$	rule #1
4)	$\text{Hg}_2\text{SO}_4_{(\text{s})}$	rule #5
5)	$\text{Pb}(\text{ClO}_4)_2_{(\text{aq})}$	rule #1
6)	$\text{SrF}_2_{(\text{s})}$	rule #4
7)	$\text{MgC}_2\text{O}_4_{(\text{s})}$	rule #7
8)	$\text{Li}_3\text{PO}_4_{(\text{aq})}$	rule #1
9)	$\text{AgI}_{(\text{s})}$	rule #3
10)	$\text{Mn}(\text{OH})_4_{(\text{s})}$	rule #6

Electrolytes:

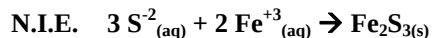
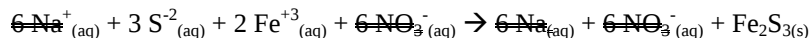
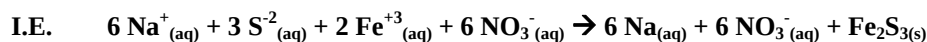
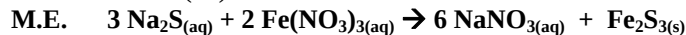
1)	$(\text{NH}_4)_2\text{S}_{(\text{aq})}$	strong electrolyte -- soluble salt
2)	$\text{HC}_7\text{H}_6\text{O}_2_{(\text{aq})}$	weak electrolyte -- weak acid
3)	$\text{N}_2\text{O}_{5(\text{l})}$	non electrolyte -- molecular compound
4)	$\text{Hg}_2\text{SO}_4_{(\text{s})}$	weak electrolyte -- insoluble salt
5)	$\text{HClO}_4_{(\text{aq})}$	strong electrolyte -- strong acid
6)	$\text{SrF}_2_{(\text{s})}$	weak electrolyte -- insoluble salt
7)	$\text{Rb}_2\text{C}_2\text{O}_4_{(\text{aq})}$	strong electrolyte -- soluble salt
8)	$\text{H}_3\text{PO}_4_{(\text{aq})}$	weak electrolyte -- weak acid
9)	$\text{AgI}_{(\text{s})}$	weak electrolyte -- insoluble salt
10)	$\text{Sr}(\text{OH})_2_{(\text{aq})}$	strong electrolyte -- soluble salt

Net ionic equations:

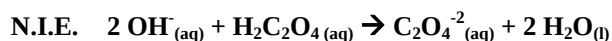
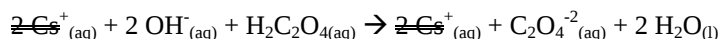
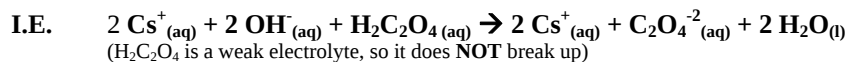
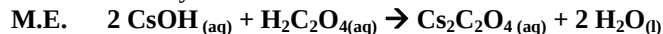
potassium phosphate and mercury (I) acetate



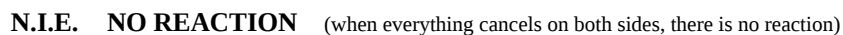
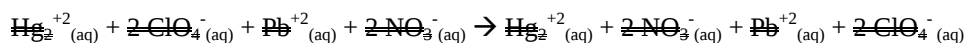
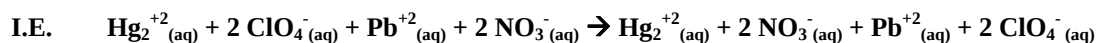
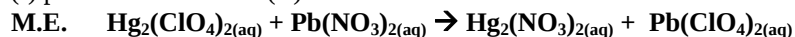
sodium sulfide with iron (III) nitrate



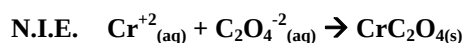
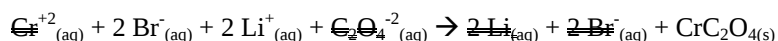
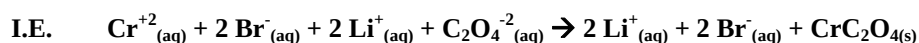
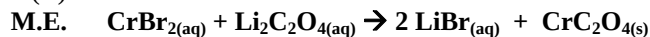
oxalic acid and cesium hydroxide



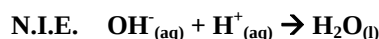
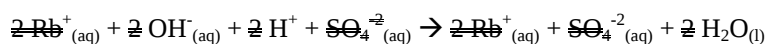
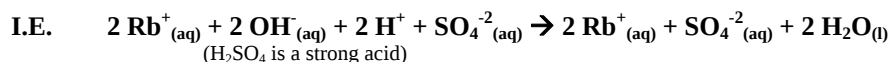
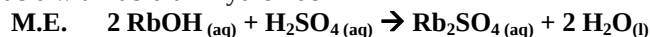
mercury (I) perchlorate and lead (II) nitrate



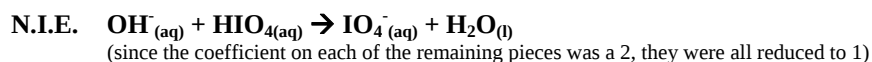
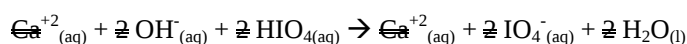
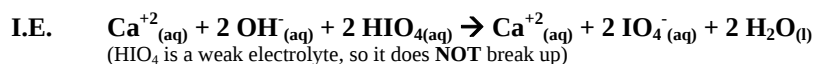
chromium (II) bromide with lithium oxalate



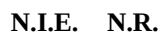
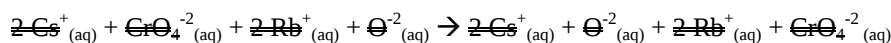
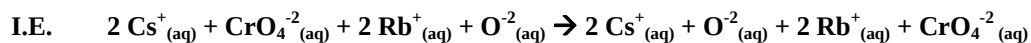
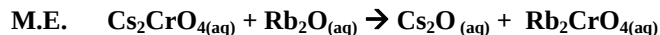
sulfuric acid with rubidium hydroxide



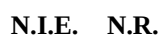
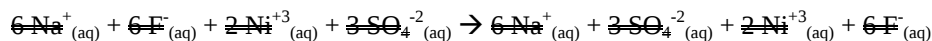
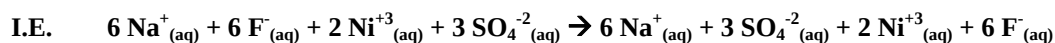
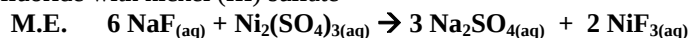
calcium hydroxide with periodic acid



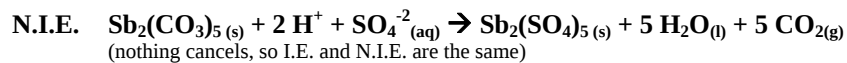
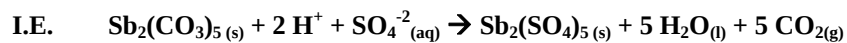
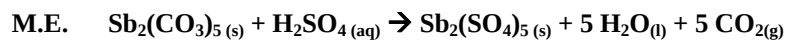
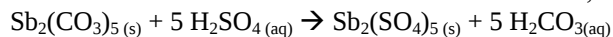
cesium chromate with rubidium oxide



sodium fluoride with nickel (III) sulfate



antimony (V) carbonate reacts with sulfuric acid to form carbon dioxide, water, and antimony (V) sulfate



lithium acetate with copper (II) nitrate

