

Name: _____

Chem. 4 – Fall 2009 – Proficiency Exam #1 – Nomenclature (1st try)

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|---|---|
| 1) lithium phosphide
Li_3P | 2) iron (III) hydroxide
$\text{Fe}(\text{OH})_3$ |
| 3) potassium phosphite
K_3PO_3 | 4) sodium oxalate
$\text{Na}_2\text{C}_2\text{O}_4$ |
| 5) $\text{Ca}(\text{ClO}_2)_2$
calcium chlorite | 6) $\text{Co}(\text{IO}_4)_3$
cobalt (III) periodate |
| 7) cadmium acetate
$\text{Cd}(\text{C}_2\text{H}_3\text{O}_2)_2$ | 8) ammonium carbonate
$(\text{NH}_4)_2\text{CO}_3$ |
| 9) aluminum selenide
Al_2Se_3 | 10) phosphoric acid
$\text{H}_3\text{PO}_{4(\text{aq})}$ |
| 11) RbHCO_3
rubidium hydrogen carbonate
rubidium bicarbonate | 12) HBr
hydrogen bromide
hydrogen monobromide |
| 13) Zn_3N_2
zinc nitride | 14) XeF_6
xenon hexafluoride |
| 15) hypobromous acid
$\text{HBrO}_{(\text{aq})}$ | 16) copper (I) iodate
CuIO_3 |
| 17) barium sulfide
BaS | 18) titanium (IV) sulfate
$\text{Ti}(\text{SO}_4)_2$ |
| 19) AgCl
silver chloride | 20) nickel (II) permanganate
$\text{Ni}(\text{MnO}_4)_2$ |
| 21) antimony (V) nitrite
$\text{Sb}(\text{NO}_2)_5$ | 22) $\text{Sn}(\text{NO}_3)_2$
tin (II) nitrate |
| 23) dinitrogen tetroxide
N_2O_4 | 24) K_2O
potassium oxide |
| 25) $\text{Bi}_2(\text{CrO}_4)_5$
bismuth (V) chromate | 26) $\text{H}_2\text{SO}_{3(\text{aq})}$
sulfurous acid |
| 27) MnO_2 (there are 2 names for this, 1 pt extra for getting both)
manganese (II) peroxide
manganese (IV) oxide | 28) $\text{HCN}_{(\text{aq})}$
hydrocyanic acid |
| 29) $\text{Pb}_3(\text{AsO}_4)_4$
lead (IV) arsenate | 30) $\text{Cr}_2(\text{SeO}_4)_3$
chromium (III) selenate |